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CHAPTER 8 SOLUTIONS MANUAL Covalent Bonding Covalent Bonding Solutions Manual Chemistry: Matter and Change □ Chapter 8 121 Section 8.1 The Covalent Bond pages 240–247 Practice Problems page 244 Draw the Lewis structure for each molecule. 1. PH<sub>3</sub> H<sub>2</sub> H<sub>2</sub>H— H<sub>2</sub>H P respectively, for single, double, and triple P — — 2. H<sub>2</sub>S H<sub>2</sub>H H — H S S ...

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Chapter 8 Covalent Bonding and Molecular Structure 8-6 Ex ample: PF<sub>3</sub> Ex CO<sub>2</sub> Step 1: Count Valence Electrons Count the total number of valence electrons in the molecule or ion. Anions have extra electrons, so add 1 electron for each negative charge. Cations have a deficiency of electrons, so subtract 1 electron for each positive charge.

~~Chapter 8: Covalent Bonding and Molecular Structure~~

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chapter 8 test chemistry covalent bonding Flashcards and ... Section 8.2 – The Nature of Covalent Bonding In ionic bonding, atoms transfer electrons to achieve noble gas configuration. In covalent bonding, atoms share electrons to achieve noble gas configuration.

### ~~Chapter 8 Covalent Bonding Assessment Answers~~

8.8: Strengths of Covalent Bonds Bond enthalpy is the enthalpy change for the breaking of a particular bond in one mole of a gaseous substance o The designation  $D(\text{bond type})$  is use to represent bond enthalpies Many important bonds such as C—H bond only exist in polyatomic molecule

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In covalent bonding, atoms still want to achieve a noble gas configuration (the octet rule). 12. In covalent bonding, atoms still want to achieve a noble gas configuration (the octet rule). But rather than losing or gaining electrons, atoms now share an electron pair. 13. In covalent bonding, atoms still want to achieve a noble gas ...

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