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single atom with a positive  
or negative charge resulting  
from the loss or gain of  
electrons, respectively.

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Writing Formulas for Acids  
and Bases. An acid is a  
compound that produces  $H^+$   
ions when it dissolves in  
water. The formula for an  
acid normally starts with  
and H. When naming acids,  
you should first determine  
whether or not the acid  
contains oxygen.



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base. Cation. Ionic  
compounds. Compounds that  
contain one or more hydrogen  
atoms and produce... An ionic  
compound that produces  
hydroxide ions when  
dissolved... Any atom or

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### Chemical Names And

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group of atoms that has a positive charge. Compounds composed of cations and anions.

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Chapter 9 Chemical Names and Formulas 83 SECTION 9.3

NAMING AND WRITING FORMULAS FOR MOLECULAR COMPOUNDS

(pages 268–270) This section explains the rules for naming and writing formulas for binary molecular compounds. Naming Binary Molecular Compounds (pages 268–269) 1. Circle the letter of the type(s) of elements that form binary molecular compounds.

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SECTION 9.2 NAMING AND  
WRITING FORMULAS FOR IONIC  
COMPOUNDS I. Write the

formulas for these binary  
ionic compounds. c.

potassium iodide 2. write  
the formulas for the

compounds formed c. b. f. 3.

Name the following binary  
ionic compounds. sodium  
sulfide g.  $\text{CuCl}$ : h.  $\text{SnCl}_2$ .

a.  $\text{MnO}_2$  c. d.  $\text{SrBr}_2$  221

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Naming and Writing Formulas  
for Ionic Compounds 9.3

Naming and Writing Formulas  
for Molecular Compounds 9.4

Naming and Writing Formulas  
for Acids and Bases 9.5 The  
Laws Governing How Compounds  
Form

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Chapter 9 “Chemical Names  
and Formulas” ... in  
parenthesis after the name  
of the metal (Table 9.2,  
p.255) Predicting Ionic  
Charges Some of the post-  
transition elements also  
have more than one possible  
oxidation state. Tin (II) =  
 $\text{Sn}^{2+}$  Lead (II) =  $\text{Pb}^{2+}$  Tin

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(IV) = Sn<sup>4+</sup> Lead (IV) = Pb<sup>4+</sup>

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## Chemical Names And

281 CHAPTER 9 Assessment 42.

a. 2- b. 1+ c. 1- d. 3+ 43.

a. 2+ b. 2+ c. 3+ d. 1+

a. barium ion b. iodide ion

c. silver ion d. mercury(II)

ion 45. cyanide,  $\text{CN}^-$  and

hydroxide,  $\text{OH}^-$  46. a.

hydroxide ion b. lead(IV)

ion c. sulfate ion d. oxide

ion 47. zero; A compound is electrically neutral. 48.

The symbols for the cation and

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Chapter 9 Practice Problems:

Chemical Names and Formulas.

Section 9.1: Naming Ions. 1.

What is the charge on the ion typically formed by each element?

a. oxygen c. sodium

e. nickel, two electrons

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t. b. iodine d. aluminum  
f. magnesium. 2. How many  
electrons does the neutral  
atom gain or lose when each  
ion forms? a.  $\text{Cr}^{3+}$  c.  $\text{Li}^+$  e.  
 $\text{Cl}^-$  b.  $\text{P}^{3-}$  ...

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