

Chemistry Post Lab Answers

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~~Answers to Post-Lab Questions for Experiment 5. I have 40 L of methane gas measured at standard temperature and pressure (STP). How many moles of methane gas is this? 1 mole of any gas at STP = 22.4 L Moles of methane = (40 / 22.4) = 1.79 moles. What is the volume of 0.30 moles of nitrogen gas if measured at STP? 1 mole of any gas at STP = 22.4 L Volume of nitrogen = 0.3 x 22.4 = 6.72 L.~~

~~SCC1123 Experiment 5 Post Lab Questions Answers - ECU ...~~

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~~Favorite Answer. First calculate the moles of KHP (potassium hydrogen pththalate, KHC8H4O4). The molar mass for KHP = 204.2 g/mole. 0.531 g KHP x (1 mole KHP / 204.2 g KHP) = 0.00260 moles KHP. KHP...~~

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~~Favorite Answer. The white precipitate most likely came from the presence of chloride ions. Ag+ + Cl- -->AgCl(s) The chemical responsible for the iron is Fe (iron), in some form. It could be almost...~~

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~~Follow this series of questions. When you can answer " yes: " to a question, then stop. 1) Does your reaction have oxygen as one of it's reactants and carbon dioxide and water as products? Yes: it's a combustion reaction 2) Does your reaction have two (or more) chemicals combining to form one chemical? Yes: it's a synthesis reaction~~

~~Chemistry of Copper~~

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~~Chemistry 2A Lab Manual~~

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