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## ~~Review~~ **Chemistry Study Guide Solution Concentration**

The concentration of a solution can be calculated using: the mass of dissolved solute in grams,  $g$  the volume of solution (or solvent) in cubic decimetres,  $dm^3$  \

$$[\text{concentration} \sim \text{in} \sim g \sim dm^{-3}] = \frac{\dots}{\dots}$$

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## **Concentration of solution - Calculations for all students ...**

Concentration of solutions A solution forms when a solute dissolves in a solvent. The concentration of a solution is a measure of how 'crowded' the solute particles are. The more concentrated the...

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## Concentration of solutions - answers

### Calculations in chemistry ...

Study Guide – Honors Chemistry

Solubility, Concentration, Equilibrium

Equilibrium: 12) Use the equation below to answer the following questions:  $3 \text{ Fe (s)} + 4 \text{ H}_2\text{O (g)} \rightleftharpoons \text{Fe}_3\text{O}_4 \text{ (s)} + 4 \text{ H}_2 \text{ (g)} + \text{Energy}$  a) Predict the shift (left or right) if

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the amount of water is increased. \_\_\_\_\_ b)  
Predict the shift (left or right) if H

## **Study Guide – Honors Chemistry Solubility, Concentration ...**

Molarity is a way to express the concentration of a solution. Units – moles & liters  
Molarity (M) – amount of solute

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(mol)/volume of solution (L) D. If used as a conversion factor, molarity would be used to convert between what two units? Between moles of solute & liters of solvent.

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Concentration Answers Chemistry Study  
Guide Solution Concentration Because the  
molarity of is the same as the overall  
molarity of  $\text{Ag}_2\text{CO}_3$  in the solution, call  
the carbonate concentration  $x$  and the  
silver ion concentration  $2x$ . The solution,  
then, is  $0.000129\text{ M Ag}_2\text{CO}_3$ , which is  
identical to the value found for the



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## **Chemistry Study Guide Solution Concentration Answers**

The amount of a substance dissolved in a given amount of solvent is the concentration of the solute, which can be expressed in terms of molarity or molality.

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If you know the molarity of a solution, you can determine the exact volume of the solution that contains a desired amount of the solute.

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and percent mass indicate. percent mass.  
molarity. heat, saturation, agitation, nature  
of particles, surface area. concentration.  
(mass of solute/mass of solution) x 100.  
moles of solute/liters of solution.

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dissolved in a given amount of solvent is  
the concentration of the solute, which can  
be expressed in terms of molarity or  
molality. If you know the

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the carbonate concentration  $x$  and the

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silver ion concentration  $2x$ . The solution, then, is  $0.000129 \text{ M Ag}_2\text{CO}_3$ , which is identical to the value found for the concentration. ...

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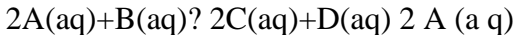
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A chemistry lab class has a solution that has a concentration of 0.391 M for A (aq) and 0.646 M for B (aq). There are no other solutes initially. The reaction



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Concentration Units. A solution is a mixture of two or more substances that is of the same composition throughout. The host substance is a solvent, and the

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dissolved substance is a solute. Although the most familiar solvents are liquids, like water or ethyl alcohol, the general concept of a solution includes solvents that are gases or even solids. In a solution, the ratio of solvent to solute is not fixed, and it can vary over a wide range, unlike compounds that are composed of definite

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solution with a high concentration of H  
ions is (very acidic/very basic). 2.  
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energy than they release are  
(endergonic/exergonic). 3.

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The concentration of a solute in this

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