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Questions | Brief  
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~~Detailed Solution~~

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~~Physical Chemistry.~~

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Problems - Molarity,

Mass Percent, and

Density of Solution

Examples Steady state



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igits.  $(31^\circ)^2$  rad

$360 = 0.54$  rad

$\sin(30.5^\circ) = 0.5075$

$\sin(31.5^\circ) = 0.5225$

$\sin(31^\circ) = 0.51$

$\cos(30.5^\circ) = 0.86163$

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$$\cos(31.5^\circ) = 0.85264$$

$$\cos(31^\circ) = 0.8615$$

Some elementary chemistry textbooks give the value of  $R$ , the ideal gas constant, as  $0.0821 \text{ l atm K}^{-1} \text{ mol}^{-1}$ .

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Write this function in

terms of Cartesian

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coordinates.  $\cos(\theta)$

$$= \frac{z}{\sqrt{x^2 + y^2 + z^2}}$$

$$= \frac{z}{\sqrt{a^2 + z^2}}$$

$$= \frac{z}{\sqrt{a^2 + z^2}}$$

$$= \frac{z}{\sqrt{a^2 + z^2}}$$

$$= \frac{z}{\sqrt{a^2 + z^2}}$$

Mathematics for

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The  $\cos(\theta)$  function is

$$\text{given by } \cos(\theta) = \frac{z}{\sqrt{a^2 + z^2}}$$

$$= \frac{z}{\sqrt{a^2 + z^2}}$$

$$= \frac{z}{\sqrt{a^2 + z^2}}$$

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e46 Mathematics for  
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where  $M$  is the molar  
mass and  $N_{Av}$  is

Avogadro ' s

constant:  $v_p = \sqrt{2mk_B T}$

$= \sqrt{2(1.3806568 \times$

$10^{-23} \text{ J}$

$\text{K}^{-1})(298.15 \text{ K})$

$(4.652 \times 10^{-26} \text{ kg})$

$1/2 = 420.7 \text{ m s}^{-1}$

For a rough graph, we

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omit a constant factor  
and plot the function  
 $g = u^2 e^{-u^2}$  where  $u$   
 $= v/v_p$ . 31.

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Suppose we have  $x=3$ ,  $y=5$  and  $a=6$  then:

$$1. a(x+y) = 6(3+5) = 6 \times 8 = 48$$
$$2. (ax) + (ay) = (6 \times 3) + (6 \times 5) = 18 + 30 = 48$$

So  $a(x+y) = (ax) + (ay)$  and the distributive rule

holds. Factorisation. Factorisation is used to tidy up equations in order to make them easier to read and understand.

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Maths for Chemists

Solutions to exercises

1B.1(b) The mean

speed is  $\sqrt{8RT/M}$

The mean

translational kinetic

energy is  $\frac{3}{2}kT$

The ratios of

species 1 to species 2

at the same

temperature are  $\frac{v_{\text{mean},1}}{v_{\text{mean},2}}$

(i)

$\frac{1}{2} \frac{m_2}{m_1}$

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$v_{\text{mean,H 1}} = 2 m v^2$

and  $2 v_{\text{mean,Hg}}$

200.6 4.003 1 2 2

$m v_{\text{rms}} E_k 1 E_k 2$

7.079 (ii) The mean

translation kinetic

energy is independent

of molecular mass

and depends upon

temperature alone!

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